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Unit 7 Worksheet 3 Questions 6, 9 done in class (check the answer file above if you missed the class) Unit 7 Worksheet 3 Hess's Law (page 2) #1 done in class Wednesday 26th February Daily Work: to be submitted by 3PM TODAY 1. Complete Hess's Law questions 2-5 (Unit 7 Worksheet 3 page 2).

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Unit 7: Thermochemistry

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Thermochemistry Quiz Review Name: _____ 1) How many joules are required to raise 20.0 g of ice from $-10.0\text{ }^{\circ}\text{C}$ to water at $50.0\text{ }^{\circ}\text{C}$? (Answer: 11274 J) 2) How much heat is required to raise the temperature of 45.6 g of copper from $11.0\text{ }^{\circ}\text{C}$ to $81.0\text{ }^{\circ}\text{C}$? The specific heat of copper is $0.38452\text{ J/g }^{\circ}\text{C}$. (Answer: 1230 J)

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Answers: 1. 4.21 g 2. -86.0 kJ/mol 3. $1.9 \times 10^2\text{ kJ}$ 4. 46 kJ/mol 5. -4521 kJ/mol The specific heat capacity of diamond is $0.5050\text{ J/g}^{\circ}\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ? A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C .

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In the table below, show how the masses,. 1 Homework - Complete and study review package - Review Appendix A5 (pp. - Homework Answers (PDF) - Sample problem answers - Thermochemistry Worksheet 1 - Thermochemistry Worksheet 2 - Worksheet 1 & 2 Answers (PDF) - Additional Sample Problems (PDF) Here are the additional practice problems we didn't get to.

Chemistry Unit 8 Worksheet 1 Answers

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Unit 7: Thermochemistry

Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words ΔH is positive? I. $\text{H}_2\text{O}(\text{l}) + 10,5\text{kcal} \rightarrow \text{H}_2\text{O}(\text{g}) \Delta H1$ II. $2\text{NH}_3 + 22\text{kcal}$

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AP Chemistry Practice Test, Ch. 6: Thermochemistry Name _____ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure. A) endothermic, positive

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