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not do "long division" to try to get exact values. Remember it is a MC test, use the answers • Mark which questions you would like to "go over" when we get to school in September. 1. Balance the following equation: $__ \text{NH}_3 + __ \text{O}_2 \rightarrow __ \text{NO}_2 + __ \text{H}_2\text{O}$ The

balanced equation shows that 1.00 mole of NH_3 requires ___ mole(s) of O_2 a. 0.57 b. 1.25 c. 1.33

Practice Test Ch 3 Stoichiometry Name Per

fewer steps are required to solve stoichiometry problems when the reactant is given in moles and the product is sought in moles which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield %yield = (act. yield/theo. yield) \times 100

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Practice Problems (Chapter 5): Stoichiometry

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 Stoichiometry Anatomy of a Chemical Equation Reactants appear on the left side of the equation. $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$...

~~Chapter 3 Stoichiometry – Chemistry~~

Stoichiometry is the tool for answering these questions.
 Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

~~Chapter 11: Stoichiometry~~

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~~Sample Exam Questions~~

Answer outline and marking scheme for question: 2. Molar mass of $\text{HgO} = 201 + 16 = 217 \text{ g mol}^{-1}$. 1.08g of HgO will contain $1.08 / 217 \text{ mols} = 0.005 \text{ mol}$ From the equation, 1 mole of $\text{O}_2(\text{g})$ is produced from 2 moles of HgO This means that 0.005 mol of HgO will produce $0.005 / 2 \text{ mol of O}_2 = 0.0025 \text{ mol}$ 0.0025 mol of O_2 will occupy $0.0025 \times 24 \text{ dm}^3 = 0.06 \text{ dm}^3$

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 fewer steps are required to solve stoichiometry problems when the reactant is given in moles and the product is sought in moles which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield $\% \text{ yield} = (\text{act. yield} / \text{theo. yield}) \times 100$

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/ 217 mols = 0.005mol From the equation, 1 mole of O₂(g) is produced from 2 moles of HgO This means that 0.005 mol of HgO will produce 0.005 / 2 mol of O₂ = 0.0025 mol 0.0025 mol of O₂ will occupy 0.0025 x 24dm³ = 0.06dm³

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