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Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used

and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass.

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We can see from the stoichiometry of the reaction that  $3/2$  mol of  $O_2$  is required to produce 1 mol of  $H_2SO_4$ . This is a standard stoichiometry problem of the type presented in Section 11.4, except this problem asks for the volume of one of the reactants ( $O_2$ ) rather than its mass. We proceed exactly as in Section 11.4, using the strategy

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### CHAPTER 11: STOICHIOMETRY

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### Chapter 11: Stoichiometry

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volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

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Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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Stoichiometry Stoichiometry CHAPTER 11 SOLUTIONS MANUAL Section 11.1  
Defining Stoichiometry pages 368–372  
Practice Problems pages 371–372 1.  
Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.  
 $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$  1 ...

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Perform a mass-to-mass stoichiometric calculation between the two reactant, using the limiting reactant ( $Cl_2$ ) as the known quantity and the excess reactant ( $S_8$ ) as the unknown quantity.  $100g Cl_2 \times 1mol Cl_2 / 70.91g Cl_2 \times 1 mol S_8 / 4mol Cl_2 \times 265.5g S_8 = 93.6g S_8$

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